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SOLUTION: How do I solve: The absolute value of $n-6$ = the ...

How many electrons are in a $n=6$ orbital with $l=2$ and $m_l=-1$? I know the total number possible is 72, I just don't know how to get it when they specify l and m_l so if someone

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could show me how to
get the answer I'd
really appreciate it!

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students will determine
how many orbitals in
any given atom can
have the following
quantum numbers or
designation

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$-(n - 6) = -(1 - n)$ and
the minus signs on
both sides change this
to $-n + 6 = -1 + n$
subtract n from both
sides and you get $-2n$
 $+ 6 = -1$ subtract 6
from both sides and
you get

**How many electrons
are in a $n=6$ orbital
with $l=2$ and $m_l=-1$**

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 $6(n-1)=2(n+1)$ Simplify
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Make love like the first
time Get high on my
own supply Dream on
or the dream dies
that's the move Cuz I

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can see the shape shift
And I feel the weight
lift Someday someday
Some way some way

Quia - how many orbitals?

with doug collins, ben
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johnson, lakewood
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Solve for n

$$6(n-1)=2(n+1) |$$

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