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...

A metal crystallizes with a face-centered cubic lattice. The edge of the unit cell is  $408 \text{ pm}$ . The diameter of the metal atom is [CBSE AIPMT 2012]

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In a solid 'AB' having NaCl structure, 'A' atoms occupy the corners of the cubic unit cell. If all the face-centred atoms along one of the axes are removed, then the resultant stoichiometry of the solid is [IIT SCREENING 2001]