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How many moles of NaCl will be produced in this reaction? (Remember—you must base this answer on how many moles of the limiting reactant that reacted.) 1.5 mol. 0.75 mol of MgCl₂ is the limiting reactant, so we use the 6:3 ratio of NaCl to MgCl₂ and multiply $(0.75)(6/3) = 1.5$ mol.

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Calculate the average atomic masses. Round all answers to two decimal places. Record your answer with the correct units. 1. Iodine is 80% 127.1 and 17% 126.1, and 3% 128.1. Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that have the following

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ChemQuest 35 Intro, to chemical Reactions Date: Hour: Formula Review 1. Write the formulas for the following compounds: 2. Write the names for the following ...

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chemquest_69_answers - ChemQuest 69 Name Date Hour Introduction Questions 1 Now we're going to get ready to learn about pH We're going to practice taking

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A solution is made so that it is 2.5 M $\text{Ca}(\text{NO}_3)_2$. Therefore the concentration of Ca^{2+} is 2.5 M and

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the concentration of NO_3^- is 5.0 M. The total concentration of all particles is 7.5 M. Explain. Since each mole of $\text{Ca}(\text{NO}_3)_2$ breaks up into 1 mole of Ca^{2+} , the concentration of Ca^{2+} would be the same as the concentration of $\text{Ca}(\text{NO}_3)_2$.

ChemQuest 39-41 Answers - Yumpu

118 ChemQuest 39 Name: _____ Date: _____ Hour: _____ Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little. Critical Thinking Questions 1.

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