

## Chapter 14 Chemical Equilibrium

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### dynamic equilibrium Chapter 14: requirements Chemical ...

AP Chapter 14: Chemical Equilibrium & K<sub>sp</sub> 3 •Warm-ups and problems will be collected before you take the test. •Read Chapter 14: Chemical Equilibrium and Chapter 16.6-16.8: Solubility Product Answer the following problems in the space provided. For problems involving an equation, carry out the

### Chapter 14. CHEMICAL EQUILIBRIUM | pdf Book Manual Free ...

EQUILIBRIUM 185 UNIT 7 EQUILIBRIUM Chemical equilibria are important in numerous biological and environmental processes. For example, equilibria involving O<sub>2</sub> molecules and the protein hemoglobin play a crucial role in the transport and delivery of O<sub>2</sub> from our lungs to our muscles.

### Chapter 14. Chemical Equilibrium

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### AP Chapter 14: Chemical Equilibrium & K<sub>sp</sub>

Chem 1120 - Chapter 14: Chemical Equilibrium Practice Quiz 1. 1. The value of the equilibrium constant for the reaction: 2 HBr(g)  $\rightleftharpoons$  H<sub>2</sub>(g) + Br<sub>2</sub>(g) is K<sub>c</sub> = 1.26 x 10<sup>-12</sup> at 500 K. This implies that: a) the product concentrations will be large relative to the reactants at equilibrium. b) the reaction has a large negative  $\Delta G^\circ$ .

### CHAPTER 14 AND CHAPTER 15 (олоороони CHEMICAL EQUILIBRIUM ...

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5. A state of dynamic chemical equilibrium is established, such as, H<sub>2</sub>(g) + I<sub>2</sub>(g)  $\rightleftharpoons$  2HI(g); Equilibrium is the exact balancing of two opposing processes. It is a dynamic state, where both forward and reverse processes continue with equal rates. At equilibrium, the concentrations of all species in the system remain constant.

### Chemical Equilibrium - kau

The K<sub>p</sub> for this reaction is 2.14 x 10<sup>-2</sup> at about 540 K. Under one set of equilibrium conditions, the partial pressure of ammonia is P(NH<sub>3</sub>) = 0.454 atm, that of hydrogen is P(H<sub>2</sub>) = 2.319 atm, and that of nitrogen is P(N<sub>2</sub>) = 0.773 atm. If an additional 1 atm of hydrogen is added to the reactor to give P(H<sub>2</sub>) = 3.319 atm, how will the system respond? Because the stress is an increase in P ...

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Chapter 14: Chemical Equilibrium Chemical Equilibrium What does it mean to describe a chemical reaction as being in a state of dynamic equilibrium? What are the characteristics and requirements of dynamic equilibrium? What does the equilibrium constant, K represent? How can we determine (quantitatively) the composition of a reaction mixture when it is at a

### CHAPTER 14: CHEMICAL EQUILIBRIUM

Chapter 14. Chemical Equilibrium What we will learn: • Concept equilibrium • Equilibrium constant • Writing ... chemical equilibrium • Meaning of equilibrium constant • Factors that affect chemical equilibrium. GCh14-2 Chemical equilibrium A dynamic stage of a chemical reaction where the concentrations of reactants and products are ...

### Introduction Chapter 14: Chemical Equilibrium - YouTube

The Equilibrium Constant K<sub>c</sub> aA + bB  $\rightleftharpoons$  cC + dD For a general reversible reaction: K<sub>c</sub> = [A]<sup>a</sup>[B]<sup>b</sup> / [C]<sup>c</sup>[D]<sup>d</sup> Equilibrium constant expression Equilibrium constant Reactants Products For the following reaction: N<sub>2</sub>O<sub>4</sub>(g)  $\rightleftharpoons$  2 NO<sub>2</sub>(g) [N<sub>2</sub>O<sub>4</sub>] / [NO<sub>2</sub>]<sup>2</sup> K = 4.64 x 10<sup>-3</sup> (at 25 °C) c = Equilibrium equation:

### Chapter 14 Chemical Equilibrium

Chapter 14 Equilibrium Notes page 1 of 6 Chapter 14. CHEMICAL EQUILIBRIUM 14.1 THE CONCEPT OF EQUILIBRIUM AND THE EQUILIBRIUM CONSTANT Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical equilibrium: A state in which the rates of the forward and reverse reactions ...

### Chapter 14 - CHEMICAL EQUILIBRIUM

CHEMICAL EQUILIBRIUM 14.1 THE CONCEPT OF EQUILIBRIUM AND THE EQUILIBRIUM CONSTANT Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical equilibrium:

A state in which the rates of the forward and reverse reactions are equal and the concentrations of the reactants and products remain constant ...

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Example 14.2 Manipulating the Equilibrium Constant to Reflect Changes in the Chemical Equation Consider the chemical equation and equilibrium constant for the synthesis of ammonia at 25 °C:  $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightleftharpoons 2 \text{NH}_3(\text{g})$   $K = 3.7 \times 10^8$

**Chapter 14.5: Factors That Affect Equilibrium - Chemistry ...**

Chapter 14: Chemical Equilibrium Q1. A reaction with an equilibrium constant  $K_c = 1.5 \times 10^{21}$  would consist of which of the following at equilibrium: A) approximately equal reactants and products B) some reactants and products with reactants slightly favored C) some reactants and products with products slightly favored D) essentially all reactants

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☐☐ Correct answer to the question CHAPTER 14 AND CHAPTER 15 (олоороони CHEMICAL EQUILIBRIUM Ex. 1 Write the expressions for the equilibrium constants  $K_p$  of the following thermal decomposition reactions: 1. (a)  $2\text{NaHCO}_3(\text{s}) \rightleftharpoons \text{Na}_2\text{CO}_3(\text{s}) + \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$  ( - e-eduanswers.com

**Chem 1120 - Chapter 14: Chemical Equilibrium**

Calculating an Equilibrium Constant from Equilibrium Concentrations. We saw in the exercise in Example 6 in Section 14.2 that the equilibrium constant for the decomposition of  $\text{CaCO}_3(\text{s})$  to  $\text{CaO}(\text{s})$  and  $\text{CO}_2(\text{g})$  is  $K = [\text{CO}_2]$ . At 800°C, the concentration of  $\text{CO}_2$  in equilibrium with solid  $\text{CaCO}_3$  and  $\text{CaO}$  is  $2.5 \times 10^{-3}$  M. Thus  $K$  at 800°C is  $2.5 \times 10^{-3}$ . (Remember that equilibrium constants ...

**Chapter 14 - Chemical Equilibrium - Questions and Problems ...**

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**Chapter 14. CHEMICAL EQUILIBRIUM**

Chapter 14 Page 2 B. The Equilibrium Constant and the Reaction Quotient. (Section 14.2) 1. Consider the total reaction: 2. We define the Reaction Quotient  $Q_c$  for this reaction as: 3. Before equilibrium is reached,  $Q_c$  will be in time, as concentrations change. 4. As equilibrium is achieved,  $Q_c$  converges to a constant and then remains fixed.

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**Chapter 14.3: Solving Equilibrium Problems - Chemistry ...**

Equilibrium is a state in which there are no observable changes as time goes by. Chemical equilibrium is achieved when: • the rates of the forward and reverse reactions are equal and • the concentrations of the reactants and products remain constant Physical equilibrium  $\text{H}_2\text{O}(\text{l})$  Chemical equilibrium  $\text{N}_2\text{O}_4(\text{g})$  14.1  $\text{H}_2\text{O}(\text{g})$   $2\text{NO}_2(\text{g})$